



SILVER OAK UNIVERSITY

Silver college of Pharmacy (067)

Programme Name: B.Pharm (18)

Subject Name: Physical Pharmaceutics-I

Subject Code: 1180673202

Semester: III

Prerequisite:

The course deals with the various physical and physicochemical properties, and principles involved in dosage forms/formulations. Theory and practical components of the subject help the student to get a better insight into various areas of formulation research and development, and stability studies of pharmaceutical dosage forms.

Objective: Upon completion of this course the student should be able to:

Upon the completion of the course student shall be able to

1. Understand various physicochemical properties of drug molecules in the designing the dosage forms
2. Know the principles of chemical kinetics & to use them for stability testing and determination of expiry date of formulations
3. Demonstrate use of physicochemical properties in the formulation development and evaluation of dosage forms

Teaching Scheme:

| Teaching Scheme | | | | |
|-----------------|---|---|---------------|--------|
| L | T | P | Contact Hours | Credit |
| 3 | 1 | 4 | 8 | 6 |

Content:

| Unit No. | Contents | Teaching Hours | Weightage % |
|----------|---|----------------|-------------|
| 1 | Solubility of drugs: Solubility expressions, mechanisms of solute solvent interactions, ideal solubility parameters, solvation & association, quantitative approach to the factors influencing solubility of drugs, diffusion principles in biological systems. Solubility of gas in liquids, solubility of liquids in liquids, (Binary solutions, ideal solutions) Raoult's law, real solutions. Partially miscible liquids, Critical solution temperature and applications. Distribution law, its limitations and applications | 10 Hrs | 22% |
| 2 | States of Matter and properties of matter: State of matter, changes in the state of matter, latent heats, | 10 Hrs | 22% |

| | | | |
|---|---|--------|--------|
| | vapour pressure, sublimation critical point, eutectic mixtures, gases, aerosols– inhalers, relative humidity, liquid complexes, liquid crystals, glassy states, solid-crystalline, amorphous & polymorphism. Physicochemical properties of drug molecules: Refractive index, optical rotation, dielectric constant, dipole moment, dissociation constant, determinations and applications | | |
| 3 | Surface and interfacial phenomenon: Liquid interface, surface & interfacial tensions, surface free energy, measurement of surface & interfacial tensions, spreading coefficient, adsorption at liquid interfaces, surface active agents, HLB Scale, solubilisation, detergency, adsorption at solid interface. | 08 Hrs | 17.77% |
| 4 | Complexation and protein binding: Introduction, Classification of Complexation, Applications, methods of analysis, protein binding, Complexation and drug action, crystalline structures of complexes and thermodynamic treatment of stability constants. | 08 Hrs | 17.77% |
| 5 | pH, buffers and Isotonic solutions: Sorensen's pH scale, pH determination (electrometric and calorimetric), applications of buffers, buffer equation, buffer capacity, buffers in pharmaceutical and biological systems, buffered isotonic solutions. | 07 Hrs | 16% |
| | Total | 45 Hrs | 100% |

Course Outcome: After Completion of Syllabus Students will able to

| Sr. No. | CO statement | Unit No |
|---------|--|---------|
| CO-1 | State the physicochemical, properties of drug molecules, pH, and solubility. | 1 |
| CO-2 | Explain the role of surfactants, interfacial, phenomenon and thermodynamics | 2 |
| CO-3 | Describe the flow behavior of fluids and concept of complexation | 3 |
| CO-4 | Analyze the chemical stability, tests of various drug products | 4 |
| CO-5 | Understand the physical, properties of solutions, buffers, isotonicity, disperse systems and rheology. | 5 |

Teaching & Learning Methodology: -

The various methods or tools follows by the faculties to teach the above subject are:

1. Student centered learning
2. Experimental learning
3. Presentation learning

List of Tutorials/Experiments:

1. Determination the solubility of drug at room temperature
2. Determination of pKa value by Half Neutralization/ Henderson Hasselbalch equation.
3. Determination of Partition co- efficient of benzoic acid in benzene and water
4. Determination of Partition co- efficient of Iodine in CCl₄ and water
5. Determination of % composition of NaCl in a solution using phenol-water system by CST method
6. Determination of surface tension of given liquids by drop count and drop weight method
7. Determination of HLB number of a surfactant by saponification method
8. Determination of Freundlich and Langmuir constants using activated char coal
9. Determination of critical micellar concentration of surfactants
10. Determination of stability constant and donor acceptor ratio of PABA-Caffeine complex by solubility method
11. Determination of stability constant and donor acceptor ratio of Cupric-Glycine complex by pH titration method

Books Recommended: - (Latest Editions)

1. Alfred Martin, Physical Pharmacy
2. Eugene, Parott. Experimental Pharmaceutics
3. Cooper and Gunn. Tutorial Pharmacy
4. Stocklosam J. Pharmaceutical Calculations, Lea & Febiger, Philadelphia.
5. Liberman H.A, Lachman C., Pharmaceutical Dosage forms, Tablets, Volume-1 to 3, MarcelDekkar Inc.
6. Liberman H.A, Lachman C, Pharmaceutical Dosage forms. Disperse systems, volume 1, 2, 3. Marcel Dekkar Inc.
7. Ramasamy C and ManavalanR, Physical Pharmaceutics.
8. C.V.S. Subramanyam, J. Thimma settee Laboratory Manual of Physical Pharmaceutics
9. C.V.S. Subramanyam, Physical Pharmaceutics
10. Gaurav Jain & Roop K. Khar Test book of Physical Pharmacy, List of Open Source

CO-PO-PSO Matrix:

| Co. No. | PO1 | PO2 | PO3 | PO4 | PO5 | PO6 | PO7 | PO8 | PO9 | PO10 | PO11 | PSO1 | PSO2 |
|----------------|------------|------------|------------|------------|------------|------------|------------|------------|------------|-------------|-------------|-------------|-------------|
| CO-1 | 3 | 2 | 3 | 2 | 3 | 2 | 1 | 2 | 2 | 3 | 2 | 3 | 3 |
| CO-2 | 3 | 2 | 3 | 3 | 3 | 2 | 2 | 3 | 3 | 2 | 3 | 2 | 3 |
| CO-3 | 3 | 2 | 2 | 3 | 2 | 3 | 2 | 2 | 2 | 3 | 2 | 3 | 3 |
| CO-4 | 3 | 2 | 3 | 2 | 2 | 2 | 3 | 2 | 3 | 2 | 2 | 2 | 3 |
| CO-5 | 2 | 3 | 2 | 3 | 2 | 2 | 3 | 2 | 3 | 2 | 3 | 3 | 3 |